

What are alkali metals and their properties?

What are alkaline earth metals and their properties?

What are halogens and their properties?

What are noble gases and their properties?

Give the symbol for the elements with the following configurations

____[Ne] $3s^23p^1$

____[Ne] $3s^23p^5$

____[Ar] $4s^2$

____[Ne] $3s^1$

Use the elements listed below to answer the following questions

K Ca Na Li Ne F Cl C Ti

____ Has the lowest ionization energy

____ Has the smallest atomic radius

____ Would most strongly attract electrons in a compound

____ Would need to absorb the most amount of energy to take on an electron

____ Is a noble gas

____ Is a transition metal

____ Has the lowest electronegativity

____ Is the largest atom

Of the pairs below pick which larger

____ Na or Ca

____ Li or F

____ H or H^+

____ Ar or Kr

____ Li or Fr

____ Li or Cs

____ Ca or Mg

____ Si or Pb

____ He or H

How does nuclear charge effect atomic radius, ionization energy, and electronegativity?